

Cite this: *RSC Adv.*, 2019, 9, 460Received 23rd October 2018
Accepted 18th December 2018

DOI: 10.1039/c8ra08771g

rsc.li/rsc-advances

Aggregation induced emission by pyridinium–pyridinium interactions†

Kaspars Leduskrasts and Edgars Suna *

Non-covalent intermolecular interactions between pyridinium subunits in a crystal-state are an efficient means to accomplish aggregation induced emission and avoid aggregation caused quenching.

Introduction

Organic luminescent molecules (luminophores) have found widespread applications in the design of optoelectronic devices and sensors.¹ In the majority of applications, organic luminophores are either used as thin solid films or they are doped into a polymer matrix (host material). In thin films and in polymer matrices highly aggregated luminophore molecules may experience intermolecular π – π aromatic interactions. The non-covalent π – π interactions usually lead to decay of the excited-state energy of organic luminophores *via* non-radiative intermolecular energy transfer, thus resulting in decrease of emissive properties in the solid state. This phenomenon is well-known as aggregation caused quenching (ACQ).² Due to the ACQ, the majority of organic luminophores feature considerably reduced emission or even lack of emission in the solid state, while being highly emissive in diluted solutions.

A traditional approach to minimize the detrimental intermolecular π – π stacking relies on a decrease of planarity of organic luminophores. This has been achieved by out-of-plane twisting of aromatic subunits in luminophores as well as by incorporation of steric bulk³ or an anionic moiety⁴ in proximity to the π -system. Combined with a restriction of intramolecular motions (to minimize the non-radiative dissipation of exciton energy), these approaches have resulted in the design of organic emitters that feature higher emission in the solid-state as compared to solution (known as aggregation induced emission or AIE). First introduced by Tang,⁵ AIE luminogens (AIEgens) have become a hot topic in materials science during the last decade,⁶ and they have found ample application in materials science.⁷ It should be noted, however, that AIEgens may lose the solid state emission if intermolecular non-covalent π – π

interactions between emitter molecules can take place.⁸ Clearly, the search for a new type of interactions that would help to avoid the detrimental intermolecular π – π stacking would facilitate the development of AIEgens with improved emissive properties.

It has been demonstrated that the undesired π – π stacking can be avoided by increasing the distance between interacting π -systems in the solid state.^{8a} We hypothesized that other interactions that are effective at longer distances and are more flexible with respect to mutual spatial orientation⁹ of the interacting π -systems can be utilized to avoid the undesired π – π stacking. This has led us to propose that interactions between heteroaromatic cationic π -systems can be employed in the design of solid-state luminogens (AIEgens). Interestingly, the pyridinium cation–arene π system interactions have been used by Bull and Fossey to generate luminescent responses in solution.¹⁰ In addition, Lu and coworkers reported a fluorescence turn-on in polymer matrices *via* cation– π interactions.¹¹ However, to the best of our knowledge, interactions between two heteroaromatic cationic π -systems has not been used in the design of AIEgens. Herein we demonstrate that non-covalent intermolecular interactions between two neighboring pyridinium subunits in solid-state is an efficient means to accomplish the solid-state emission (Fig. 1).

Results and discussion

Synthesis of quaternary pyridinium salts 1–5

N-Benzyl pyridinium bromide **1** was obtained by alkylation of pyridine with benzyl bromide (Fig. 2). The synthesis of pyridinium carbazoles **2a** and **2b** involved reduction of aldehyde **6** to benzylic alcohol **7**, followed by conversion to benzyl chloride and reaction with pyridine (Fig. 2). The corresponding tosylate **2b** was prepared from benzyl alcohol **7** and TsCl in the presence of pyridine as a base. In addition, pyridinium carbazoles **3**, **4** possessing substituents in position 4 of the pyridine ring were also synthesized to evaluate influence of steric bulk on the interaction between pyridinium subunits of neighboring molecules in the solid-state (Fig. 2). Finally, 1,4-butylene linker-

Latvian Institute of Organic Synthesis, Aizkraukles 21, LV-1006, Riga, Latvia. E-mail: edgars@osi.lv

† Electronic supplementary information (ESI) available: Experimental procedures, photophysical properties, X-ray crystallographic data (CIF files), ¹H and ¹³C NMR data. CCDC 1873711–1873715. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c8ra08771g



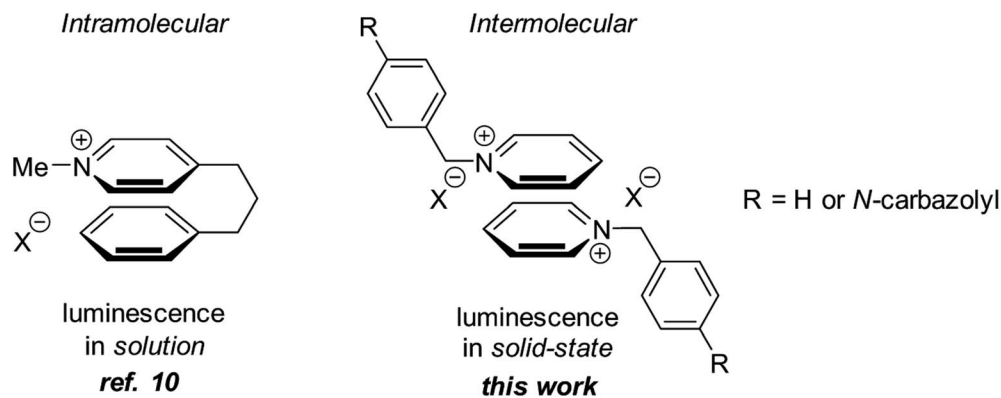


Fig. 1 Design of the solid-state AIEgens based on pyridinium-pyridinium interactions.

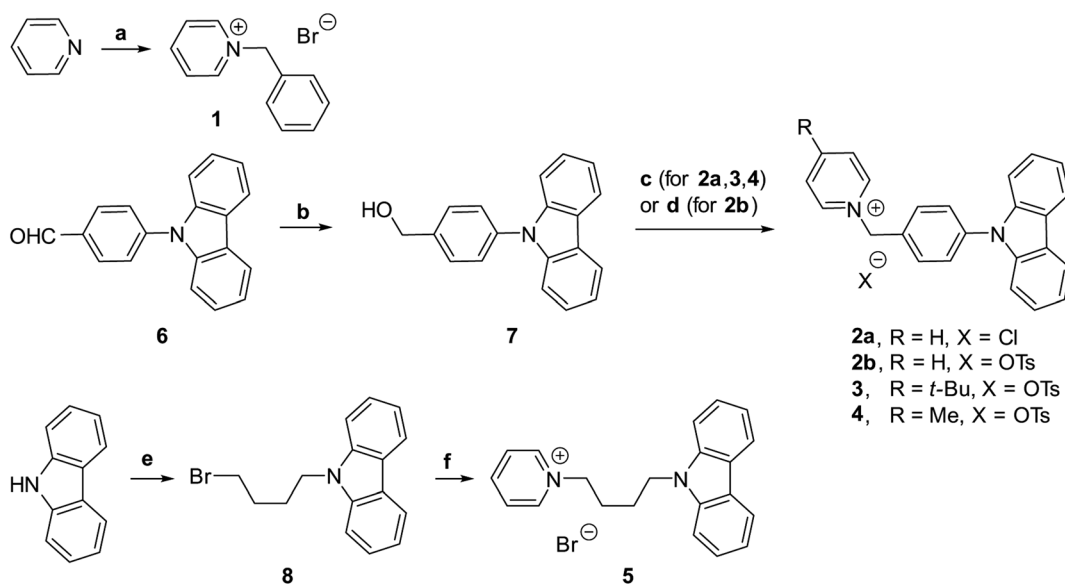


Fig. 2 Synthesis of pyridinium salts 1–5. Reagents and conditions: (a) BnBr, acetone, reflux, 1 h, 94%. (b) NaBH₄, THF, MeOH, rt, 16 h, 87%. (c) TsCl, pyridine (for 2a), 4-*t*Bu-pyridine (for 3) or 4-Me-pyridine (for 4), CH₂Cl₂, rt, 24 h, 57% (2a), 63% (3), 53% (4). (d) SOCl₂, CH₂Cl₂, rt, 30 min, then pyridine, MeCN, 80 °C, 1 h, 70%. (e) Br(CH₂)₄Br, KOH, DMSO, rt, ultrasound, 30 min, 48%. (f) pyridine, MeCN, 90 °C, 16 h, 69%.

containing pyridinium carbazole **5** was prepared by double substitution of 1,4-dibromobutane with carbazole and pyridine (Fig. 2). All obtained pyridinium salts 1–5 were crystalline materials and they were characterized by ¹H and ¹³C NMR spectra as well as by HRMS and IR techniques.

Luminescence data

UV-vis absorption spectra of 1–5 were measured in MeCN solutions (at ca. 10⁻⁶ mol L⁻¹ concentration) at room temperature and under ambient atmosphere. All synthesized materials displayed absorption peaks in 235–360 nm range (Fig. 3A), which were attributed to the π-π* transitions. Pyridinium salt **1** showed a narrow absorption band in MeCN solution with a maximum at 285 nm (entry 1, Table 1). Intense solid-state absorbance in the 250–350 nm region with sharply decreasing absorbance intensity above 350 nm was observed for the crystalline salt **1** (see ESI, page S10†). The salt also displayed an emission in MeCN solution (see

ESI, page S9†) with photoluminescence quantum yield (PLQY) of 1.9% (entry 1, Table 1). Notably, excitation of the crystalline pyridinium salt **1** at 393 nm resulted in a solid-state emission with PLQY of 5.3% (entry 1, Table 1) and the maximum at 520 nm (see ESI, page S9†). The observed 2.8-fold increase of PLQY in the solid-state as compared to that in the MeCN solution pointed to the AIE properties of pyridinium salt **1**. Next, emission properties of pyridinium salts 2–5 with attached carbazole luminophore were evaluated.

All pyridinium carbazoles 2–5 were virtually non-emissive in solutions with photoluminescence quantum yields (PLQY) below 0.1%. In sharp contrast, the solid state PLQY for pyridinium salts 2–5 were measured to be in the range from 9.7 to 18.7%. Hence, pyridinium salts featured up to 187 times increase of emission in the solid state as compared to that in MeCN solution (Table 1). These results clearly indicated that pyridinium salts 2–5 possess AIE properties. Interestingly, the solid state emission intensity was not influenced by the



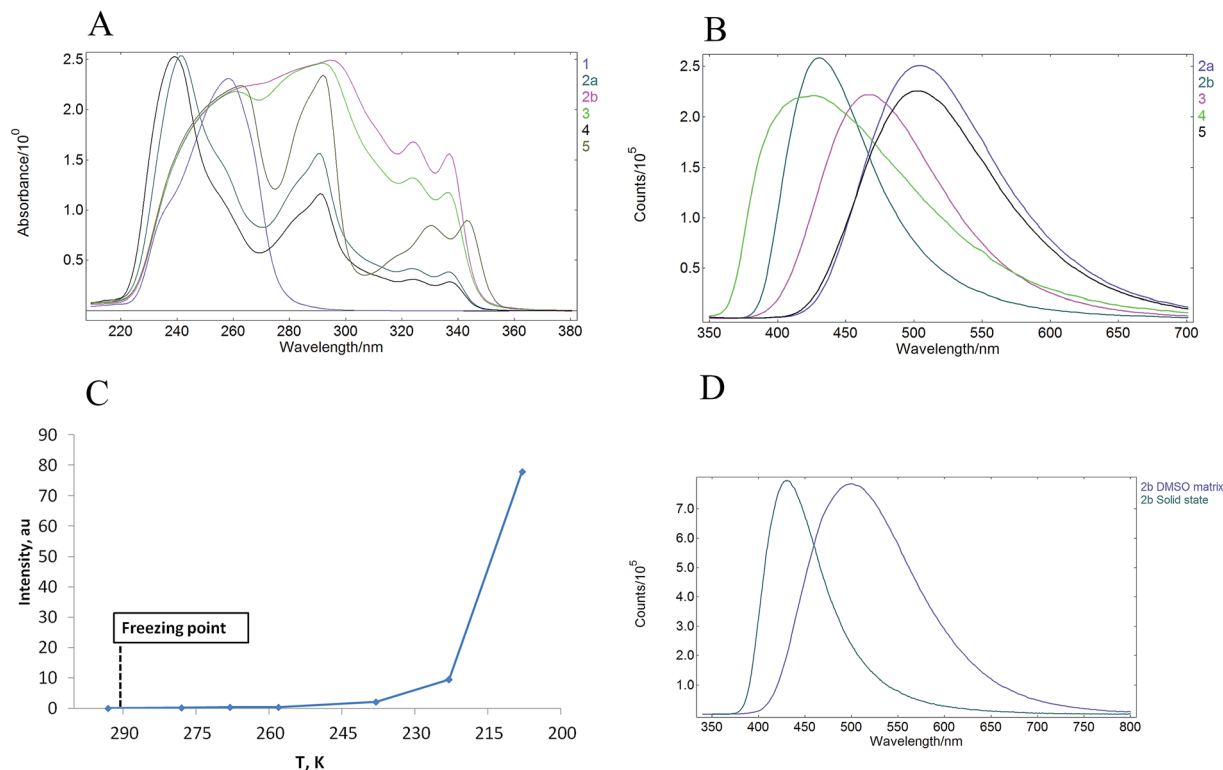


Fig. 3 (A) UV-vis spectra of 1–5 in MeCN solutions; (B) solid state emission of 2–5; (C) emission of 2b in DMSO as a function of temperature; (D) normalized intensity of emission of 2b in the solid state and in the DMSO matrix.

Table 1 Photoluminescent properties of pyridinium salts 1–5

Entry	Compound	λ_{Abs} , nm	Solid λ_{Em} , nm	Solution ϕ (%)	Solid ϕ (%)	α_{AIE}
1	1	258	520 ^a	1.9	5.3	2.8
2	2a	241, 290, 325, 337	505 ^b	<0.1	17.3	>173
3	2b	264, 295, 324, 337	432 ^b	<0.1	18.7	>187
4	3	260, 294, 325, 337	465 ^b	<0.1	18.5	>185
5	4	240, 292, 324, 337	442 ^b	<0.1	9.7	>97
6	5	263, 292, 330, 343	503 ^b	<0.1	15.9	>159

^a Excited at 393 nm. ^b Excited at 325 nm.

structure of the anion as evidenced by comparable solid state PLQY for pyridinium chloride **2a** (entry 2, Table 1) and pyridinium tosylate **2b** (entry 3). Introduction of bulky *t*-Bu substituent in the position 4 of pyridinium subunit did not affect the solid state PLQY of **3** (entry 4 vs. entry 3), however the corresponding 4-methyl substituted analogue **4** showed the lowest PLQY values in the carbazole series (entry 5). Finally, AIEgen **5** with a flexible butylene linker between carbazole subunit and pyridinium moiety featured only slightly reduced solid state PLQY as compared to **2b** (entry 6 vs. entry 3). This is notable given a higher degree of freedom for intramolecular motions in **5** as compared to **2b** due to the long alkyl chain, which may consume the excited-state through non-radiative relaxation. All AIEgens 2–5 showed a broad emission peak in the solid state with no distinctive bands except the maximum in the range of 432–520 nm (Fig. 3B). Importantly, wavelength of

the solid-state emission maximum depended on the counterion in pyridinium salts 1–5. Higher energy emission peaks in the range of 432–465 nm were observed for AIEgens **2b**, **3** and **4**, which possess tosylate as the counter ion (Table 1, entries 3–5). In contrast, lower energy emission (503–520 nm) was measured for chloride or bromide counterion-containing AIEgens **1**, **2a** and **5** (Table 1, entries 1, 2 and 6). Finally, pyridinium carbazoles 2–5 displayed intense solid-state absorbance in the 250–350 nm region with sharply decreasing absorbance intensity above 350 nm. The solid-state absorbance of 2–5 was consistent with that in MeCN solution (see ESI, page S10[†]).

Photoluminescence behaviour of **2a** in various solvent mixtures was investigated to get an additional insight in AIE properties of **2a**. Lack of the emission was observed for **2a** in acetonitrile (a good solvent). In contrast, enhanced emission was observed in mixtures containing high (90–95% v/v) fraction



of Et₂O (a poor solvent) in MeCN (see ESI, page S13†). The dynamic light scattering (DLS) measurements indicated that pyridinium salt **2a** did not form aggregates in solvent mixtures that contained less than 90% of Et₂O in acetonitrile (see ESI, page S13†). Notably, the lack of emission was observed for **2a** in these solvent mixtures. In contrast, the formation of aggregates could be observed even by the naked eye if 90–95% (v/v) of Et₂O in acetonitrile was used as the solvent. In the latter case, the size of aggregates was larger than 10 μm, thus exceeding limits of DLS measurements. The apparent correlation between the formation of aggregates and luminescence provided an additional support for AIE properties of pyridinium salt **2a**.

We have also examined whether AIE properties of pyridinium salts **1–5** originate from restriction of intramolecular rotation (RIR) phenomena.¹² To this end we measured a relationship between emission of **2b** in DMSO solution and temperature in the range from 293 K to 208 K (Fig. 3C). A solution of **2b** in DMSO did not display observable emission at 293 K. Importantly, the emission increase could not be determined after freezing of the DMSO solution of **2b**, as well as upon further cooling to 238 K. The emission only started to appear at temperatures below 238 K and the intensity increased considerably at 208 K (Fig. 3C). Lack of the emission for **2b** in the DMSO matrix points against the RIR mechanism as the origin of AIE properties of **2b**, because considerable enhancement of the emission in solid matrices is usually observed for most of AIEgens that benefit from RIR effect.¹³ Furthermore, the emission peak for **2b** in frozen DMSO matrix showed bathochromic shift as compared to that in the solid state (Fig. 3D). Finally, the solid state emission intensity for crystalline **2b** remained unchanged in the temperature range from 293 K to 193 K. These data provide evidence that the luminogens **1–5** show AIE properties in crystalline form.

Single crystal X-ray analysis

Single crystal X-ray analysis of luminogens **2–5** provided important insight into intermolecular interactions that result in AIE properties (Table 2). Single crystals of pyridinium salts **2–5** suitable for X-ray crystallography could be obtained by vapor diffusion from various solvents systems such as acetone/EtOH, DCM/hexane, EtOAc/MeCN, toluene/DCM or Et₂O. In a crystal lattice pyridinium salts **2a**, **2b**, **3** and **5** feature intermolecular $\pi^+-\pi^+$ interactions between the charged pyridinium rings. This is evidenced by a strictly parallel off-center orientation of neighboring pyridinium rings with the distance between them spanning a range from 3.686 Å (for **5**) to 4.211 Å (for **2a**; see Table 2). In contrast, pyridinium salt **4** does not have the $\pi^+-\pi^+$ interactions between the charged pyridinium rings. Instead, interactions between pyridinium ring and 1,4-disubstituted phenylene moiety can be observed for **4** as evidenced by a relatively close distance (from 3.803 Å to 4.027 Å) between the two nonparallel-oriented π -systems (Table 2).¹⁴ It should be noted that pyridinium salt **4** demonstrated lower PLQY as compared to other AIEgens **2a**, **2b**, **3** and **5** (see Table 1), suggesting that pyridinium–pyridinium interactions are important to achieve greater AIE properties in the crystalline state.¹⁵

In addition to the intermolecular pyridinium–pyridinium interactions in the crystal lattice of **2b**, the aromatic system of tosylate anion is positioned in a non-parallel off-center orientation with respect to pyridinium ring (3.843–4.186 Å distance, see ESI, page S9†). However, the latter interaction apparently does not influence the AIE properties of **2b**, as evidenced by similar PLQY values for **2b** and the analogous AIEgen **2a**, possessing chloride as the counter-ion (see entries 2 and 3, Table 1). The other tosylate-containing AIEgens **3** and **4** do not show the face-to-face interactions between pyridinium cation and π -

Table 2 Crystal packing and expanded views of $\pi^+-\pi^+$ and $\pi^+-\pi$ interactions for pyridinium salts **2–5**

Compound 2a	Compound 2b	Compound 3	Compound 4	Compound 5



system of the tosylate anion. Hence, these data demonstrate that the crystal-state AIE effect can be achieved using chloride, bromide and tosylate as the counter-ion for pyridinium. Importantly, nature of counter-ion influences the solid-state emission maximum of AIEgens. Thus, change of the counter-ion in AIEgen 2 from tosylate (**2b**) to chloride (**2a**) resulted in a noticeable bathochromic shift of more than 70 nm (Fig. 3B.). The dependence of the solid-state emission maximum on the nature of the counter-ion simplifies the design of AIEgens with the desired emission wavelength.

Conclusions

All synthesized pyridinium salts **1–5** demonstrate AIE properties with up to 187 times emission increase in the crystal-state as compared to solution. Lack of emission for **2b** in solution and frozen DMSO matrix (solid amorphous state) speaks against the RIR effect as the origin of the AIE properties for pyridinium salt **2b**. Single crystal X-ray analyses provide clear evidence for the presence of non-covalent intermolecular pyridinium–pyridinium and pyridinium– π interactions in the crystal-state of AIEgens **2–5**. Consequently, crystal-state AIE effect can be attributed to the intermolecular $\pi^+–\pi^+$ and $\pi^+–\pi$ interactions involving pyridinium cations. Hence, non-covalent interactions between two neighboring pyridinium subunits in a crystal-state are an efficient means to accomplish AIE. The nature of counter-ion in pyridinium salts **2–5** does not affect emission efficiency (PLQY) of the crystal-state AIEgens, however, the counter-ion does influence the crystal-state emission maximum of AIEgens. The use of the non-covalent pyridinium–pyridinium or pyridinium– π interactions in the design of AIEgens is alternative and complementary approach to routinely used means to avoid the ACQ effect.^{3,4} Further studies on the design of the crystal-state AIEgens based on interactions between cationic π systems will be reported in due course.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was funded by Taiwan–Lithuania–Latvia Tripartite Cooperation Fund (Project “New materials and technologies for very-high color rendering and high sunlight spectrum resemblance OLED lighting sources”). We thank Dr A. Mishnev and Dr D. Stepanovs for X-ray crystallographic analysis and P. Dimitrijevs for DLS measurements.

References

- (a) D. Volz, *J. Photonics Energy*, 2016, **6**, 020901; (b) X.-D. Wang, O. S. Wolfbeis and R. J. Meier, *Chem. Soc. Rev.*, 2013, **42**, 7834–7869.
- (a) A. S. Klymchenko, *Acc. Chem. Res.*, 2017, **50**, 366; (b) S. A. Jenekhe and J. A. Osaheni, *Science*, 1994, **265**, 765; (c) Solid-state luminescence enhancement has been also observed: J. Shi, L. E. Aguilar Suarez, S.-J. Yoon, S. Varghese, C. Serpa, S. Y. Park, L. Lüer, D. Roca-Sanjuán, B. Milián-Medina and J. Gierschner, *J. Phys. Chem. C*, 2017, **121**, 23166–23183.
- (a) J. E. Anthony, J. S. Brooks, D. L. Eaton and S. R. Parkin, *J. Am. Chem. Soc.*, 2001, **123**, 9482–9483; (b) Z.-F. Yao, J.-Y. Wang and J. Pei, *Cryst. Growth Des.*, 2018, **18**, 7–15.
- J. Wang, X. Gu, P. Zhang, X. Huang, X. Zheng, M. Chen, H. Feng, R. T. K. Kwok, J. W. Y. Lam and B. Z. Tang, *J. Am. Chem. Soc.*, 2017, **139**, 16974–16979.
- J. Luo, Z. Xie, J. W. Y. Lam, L. Cheng, H. Chen, C. Qiu, H. S. Kwok, X. Zhan, Y. Liu, D. Zhuc and B. Z. Tang, *Chem. Commun.*, 2001, 1740–1741.
- (a) T. Butler, W. A. Morris, J. Samonina-Kosicka and C. L. Fraser, *ACS Appl. Mater. Interfaces*, 2016, **8**, 1242; (b) Z. S. Wang, B. S. Gelfand, P. C. Dong, S. Trudel and T. Baumgartner, *J. Mater. Chem. C*, 2016, **4**, 2936; (c) O. Shynkaruk, G. He, R. McDonald, M. J. Ferguson and E. Rivard, *Chem.–Eur. J.*, 2016, **22**, 248; (d) M. T. Gabr and F. C. Pigge, *RSC Adv.*, 2015, **5**, 90226; (e) W. A. Morris, T. D. Liu and C. L. Fraser, *J. Mater. Chem. C*, 2015, **3**, 352; (f) K. R. Ghosh, S. K. Saha and Z. Y. Wang, *Polym. Chem.*, 2014, **5**, 5638; (g) N. B. Shustova, A. F. Cozzolino, S. Reineke, M. Baldo and M. Dinca, *J. Am. Chem. Soc.*, 2013, **135**, 13326; (h) F. Hu, G. X. Zhang, C. Zhan, W. Zhang, Y. L. Yan, Y. S. Zhao, H. B. Fu and D. Q. Zhang, *Small*, 2015, **11**, 1335; (i) C. Wang, H. Zhang, L. Tian, W. Zhu, Y. Lan, J. Li, H. Wang, G. X. Zhang, D. Q. Zhang, S. L. Yuan and G. T. Li, *Sci. China: Chem.*, 2016, **59**, 89.
- (a) Y. Hong, J. W. Y. Lam and B. Z. Tang, *Chem. Soc. Rev.*, 2011, **40**, 5361; (b) J. Mei, N. L. C. Leung, R. T. K. Kwok, J. W. Y. Lam and B. Z. Tang, *Chem. Rev.*, 2015, **115**, 11718; (c) X. Z. Yan, H. Z. Wang, C. E. Hauke, T. R. Cook, M. Wang, M. L. Saha, Z. X. Zhou, M. M. Zhang, X. P. Li, F. H. Huang and P. J. Stang, *J. Am. Chem. Soc.*, 2015, **137**, 15276; (d) X. Z. Yan, M. Wang, T. R. Cook, M. M. Zhang, M. L. Saha, Z. X. Zhou, X. P. Li, F. H. Huang and P. J. Stang, *J. Am. Chem. Soc.*, 2016, **138**, 4580; (e) A. D. Shao, Y. S. Xie, S. J. Zhu, Z. Q. Guo, S. Q. Zhu, J. Guo, P. Shi, T. D. James, H. Tian and W. H. Zhu, *Angew. Chem., Int. Ed.*, 2015, **54**, 7275; (f) C. Wang, B. J. Xu, M. S. Li, Z. G. Chi, Y. J. Xie, Q. Q. Li and Z. Li, *Mater. Horiz.*, 2016, **3**, 220; (g) Z. L. Xie, C. J. Chen, S. D. Xu, J. Li, Y. Zhang, S. W. Liu, J. R. Xu and Z. G. Chi, *Angew. Chem., Int. Ed.*, 2015, **54**, 7181; (h) Y. T. Gao, G. X. Feng, T. Jiang, C. C. Goh, L. G. Ng, B. Liu, B. Li, L. Yang, J. L. Hua and H. Tian, *Adv. Funct. Mater.*, 2015, **25**, 2857; (i) W. Dong, T. Fei, A. Palma-Cando and U. Scherf, *Polym. Chem.*, 2014, **5**, 4048.
- (a) M. Mohan, S. Pangannaya, M. N. Satyanarayan and D. R. Trivedi, *ChemistrySelect*, 2018, **3**, 3803–3813; (b) N. Han, H. Kun, X. Yuanjing, P. Qian, Z. Zujin, H. Rongrong, C. Junwu, S. Shi-jian, Q. Anjun and B. Z. Tang, *Mater. Chem. Front.*, 2017, **1**, 1125–1129.
- M. S. Marshall, R. P. Steele, K. S. Thanthiruwatte and C. D. Sherrill, *J. Phys. Chem. A*, 2009, **113**, 13628–13632.



- 10 W. Chen, S. A. Elfeky, Y. Nonne, L. Male, K. Ahmed, C. Amiable, P. Axe, S. Yamada, T. D. James, S. D. Bull and J. S. Fossey, *Chem. Commun.*, 2011, **47**, 253–255.
- 11 J. Zhong, Z. Li, W. Guan and C. Lu, *Anal. Chem.*, 2017, **89**, 12472–12479.
- 12 J. Chen, C. C. W. Law, J. W. Y. Lam, Y. Dong, S. M. F. Lo, I. D. Williams, D. Zhu and B. Z. Tang, *Chem. Mater.*, 2003, **15**, 1535–1546.
- 13 Y. Hong, J. W. Y. Lam and B. Z. Tang, *Chem. Commun.*, 2009, 4332–4353.
- 14 Similar interactions between pyridinium cation and aromatic or heteroaromatic π -system have been observed by X-ray analysis: (a) X.-H. Jin, C. Chen, C.-X. Ren, L.-X. Cai and J. Zhang, *Chem. Commun.*, 2014, 15878–15881; (b) C. Chen, X.-H. Jin, X.-J. Zhou, L.-X. Cai, Y.-J. Zhang and J. Zhang, *J. Mater. Chem. C*, 2015, **3**, 4563–4589; (c) Y. Gu, Z. Zhao, H. Su, P. Zhang, J. Liu, G. Niu, S. Li, Z. Wang, R. T. K. Kwok, X.-L. Ni, J. Sun, A. Qin, J. W. Y. Lam and B. Z. Tang, *Chem. Sci.*, 2018, **9**, 6497–6502.
- 15 9-(4-Benzylphenyl)-9H-carbazole **10** (see ESI, page S7†), an analog of **2a,b** with benzene ring instead of the pyridinium subunit does not possess AIE properties. It features high solution-state emission (39.9% PLQY) and reduced solid-state emission (35.5% PLQY). For photophysical properties of **10**, see ESI.†

